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Covalent Bonding Worksheet

Covalent bonding occurs when two or more NON-METALS share electrons, attempting to attain a stable octet (8 outer electrons) in their outer shell for at least part of the time. Draw a Lewis dot diagram for each element listed. Circle the unpaired electrons that will be shared between the elements.

1.) H ₂ (hydrogen is diatomic)	$H + H \rightarrow H \overset{\cdot\cdot}{\underset{\cdot\cdot}{\text{H}}} \rightarrow H-H$ Single Bond
2.) F ₂ (fluorine is diatomic)	$F + F \rightarrow \overset{\cdot\cdot}{\underset{\cdot\cdot}{\text{F}}} \overset{\cdot\cdot}{\underset{\cdot\cdot}{\text{F}}} \rightarrow \overset{\cdot\cdot}{\underset{\cdot\cdot}{\text{F}}} \overset{\cdot\cdot}{\underset{\cdot\cdot}{\text{F}}}$
3.) O ₂ (oxygen is diatomic)	$O + O \rightarrow \rightarrow O=O$ Double Bond
4.) N ₂ (nitrogen is diatomic. Is this a triple bond?)	$N + N$
5.) BF ₃ (just need 3 fluorine atoms here)	$B + F$
6.) Ammonia: NH ₃ (how many hydrogen atoms are needed?)	$N + H$
7.) Carbon dioxide: CO ₂	$O + C + O$
8.) Methane: CH ₄ (careful here: 4 hydrogen atoms needed)	$C + H$
9.) Dihydrogen monoxide: (the most dangerous substance on the planet. It has killed more people than any other substance known to mankind!!)	$H + O + H$
10.) SO ₂ (hint: one pair of electrons from sulfur must be sit up for this one to work.)	$O + S + O$

Covalent Bonding Worksheet

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Directions: Draw the Lewis dot diagram for the following compounds and indicate the type of bonding that occurs in the compound.

Compound	Structure 1 (based on Lewis dot diagram)	Structure 2 (based on Lewis dot diagram)	Bond Type
H ₂			Covalent
N ₂			Covalent
O ₂			Covalent
CO ₂			Covalent
H ₂ O			Covalent
CH ₄			Covalent
HF			Covalent
Cl ₂			Covalent
Carbon monoxide			Covalent
SO ₂			Covalent
SO ₃			Covalent
NO ₂			Covalent
NO			Covalent
CO			Covalent
SiO ₂			Covalent
Si			Covalent
SiO			Covalent
NO ₂			Covalent
NO			Covalent
CO			Covalent
CO ₂			Covalent

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